

A Semiempirical Method to Detect and Correct DFT-Based Gas-Phase Errors and Its Application in Electrocatalysis

Laura P. Granda-Marulanda, Alejandra Rendón-Calle, Santiago Builes, Francesc Illas, Marc T. M. Koper,* and Federico Calle-Vallejo*

Cite This: *ACS Catal.* 2020, 10, 6900–6907

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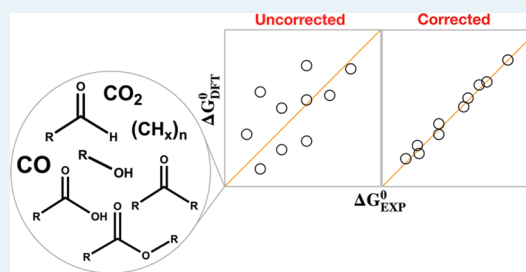
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ABSTRACT: Computational models of adsorption at metal surfaces are often based on DFT and make use of the generalized gradient approximation. This likely implies the presence of sizable errors in the gas-phase energetics. Here, we take a step closer toward chemical accuracy with a semiempirical method to correct the gas-phase energetics of PBE, PW91, RPBE, and BEEF-vdW exchange–correlation functionals. The proposed two-step method is tested on a data set of 27 gas-phase molecules belonging to the carbon cycle: first, the errors are pinpointed based on formation energies, and second, the respective corrections are sequentially applied to ensure the progressive lowering of the data set's mean and maximum errors. We illustrate the benefits of the method in electrocatalysis by a substantial improvement of the calculated equilibrium and onset potentials for CO₂ reduction to CO on Au, Ag, and Cu electrodes. This suggests that fast and systematic gas-phase corrections can be devised to augment the predictive power of computational catalysis models.

KEYWORDS: gas-phase errors, gas-phase corrections, carbon monoxide, carbon dioxide, DFT, electrochemical reduction of CO₂



INTRODUCTION

For decades, considerable effort has been devoted to increasing the accuracy of density functional theory (DFT). This has been done by developing more accurate exchange–correlation functionals at the generalized gradient approximation (GGA) level,^{1–3} hybrid functionals,^{4–6} and range-separated functionals.^{7–9} In addition, different correction schemes have been developed to account for electron localization¹⁰ or dispersion interactions.^{11–13} Lately, machine learning schemes¹⁴ have also been proposed to bypass Kohn–Sham equations. In general, these efforts include careful computational benchmarking and comparison to experiments.^{15–17}

An agreement has been reached in the scientific community about the level of theory required to simulate certain materials with a good tradeoff between computational time and accuracy. For instance, hybrid functionals are advisable for molecules and solids with localized electrons, while GGAs usually suffice for bulk and surface metals.^{13,16} However, the choice is not trivial when dealing with systems where metals and molecules are involved and ought to be simulated at the same level of theory. In such a case, the accuracy may be improved by using GGA functionals and adding semiempirical corrections to the DFT energies of molecules, as done for thermochemical reaction energies of interest in catalysis,^{18,19} formation and decomposition energies of solids,^{20,21} and catalytic kinetic barriers.^{22,23}

In this article, we provide a simple and fast procedure for detecting gas-phase errors based on the formation energies of

reactants and products calculated with DFT. Improving the description of the gas phase is shown to enhance catalytic predictive power by analyzing the electrocatalytic CO₂ reduction reaction to CO on Au, Ag, and Cu electrodes. The reduction of CO₂ and CO (hereafter denoted as CO₂RR and CORR, respectively) are of great importance in catalysis science and technology as they lead to valuable feedstocks and fuels such as methane, ethylene, ethanol, and formic acid while helping in balancing the carbon cycle.^{24–26} Although DFT has been used to predict enhanced catalysts for other electrocatalytic reactions,^{27–29} it has been, so far, challenging to elaborate robust design routines for CO₂RR and CORR to hydrocarbons and oxygenates.³⁰ Thus, the method presented here may help boost materials design via screening for those paramount reactions.

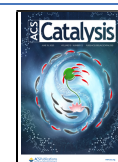
COMPUTATIONAL METHODS

All calculations were performed using the Vienna Ab initio simulation package.³¹ Dissimilar gas-phase errors have been pointed out in previous studies for the total energy of CO(g) and CO₂(g)^{18,19,32} using PBE and RPBE.³³ In addition, others

Received: March 4, 2020

Revised: May 25, 2020

Published: May 26, 2020



suggested a correction for the total energy of $\text{H}_2(\text{g})$ to be applied only when using BEEF-vdW.³⁴ Thus, we made a functional-dependent analysis including four different xc functionals habitually used in catalysis, namely, PBE,³⁵ PW91,³⁶ RPBE,³³ and BEEF-vdW.⁹ The gas-phase molecules were relaxed with the conjugate gradient algorithm in boxes of $\sim 3375 \text{ \AA}^3$, considering only the Γ point. The effect of the cores on the valence electron density is incorporated using the projector-augmented wave (PAW) method.³⁷ To compute the formation energies of the molecules, graphite was represented by graphene. Approximating graphene as the standard state of carbon is based on the weak interlayer cohesive energy of graphite ($0.031\text{--}0.064 \text{ eV/atom}$)^{38–43} (see Section S6 in the Supporting Information). The optimized interatomic distances of graphene are 1.43 (PBE and RPBE) and 1.42 Å (PW91 and BEEF-vdW).

The convergence criterion for the maximal forces on the atoms for all simulations was 0.01 eV \AA^{-1} , and the plane-wave cutoff was set to 400 eV. Convergence tests for the free energy of reaction of $\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g})$ with plane-wave cutoffs in the range of 300–1000 eV within PBE showed that 400 eV is enough to achieve accurate reaction energies with an average difference of $\sim 5 \text{ meV}$ (see Table S1). None of the species analyzed has unpaired electrons, so spin unrestricted calculations were not required. Gaussian smearing with $k_{\text{B}}T = 0.001 \text{ eV}$ was used. In all cases, the energies were extrapolated to 0 K.

The reaction free energies were obtained as $\Delta G^0 = \Delta E_{\text{DFT}} + \Delta \text{ZPE} - T\Delta S^0$ where ZPE is the zero-point energy contribution calculated from the vibrational frequencies obtained using the harmonic oscillator approximation. The standard total entropies (S^0) and the experimental standard free energies (ΔG_{exp}^0) were obtained from thermodynamic tables^{44–46} at $T = 298.15 \text{ K}$. In cases where ΔG_{exp}^0 was not tabulated, it was evaluated by combining entropy and enthalpy values: $\Delta G_{\text{exp}}^0 = \Delta H_{\text{exp}}^0 - T\Delta S_{\text{exp}}^0$. We did not include heat capacity effects as recent studies showed that formation energies are not significantly modified by them from 0 to 298.15 K.²¹

Electrocatalytic CO_2 reduction to CO was modeled based on the free energy scheme described in previous reports,⁴⁷ making use of the computational hydrogen electrode⁴⁸ for the description of proton–electron transfers. The reaction pathway proceeds via CO_2 hydrogenation (step 1: $\text{CO}_2 + \text{H}^+ + \text{e}^- + * \rightarrow * \text{COOH}$), followed by $* \text{CO}$ formation (step 2: $* \text{COOH} + \text{H}^+ + \text{e}^- \rightarrow * \text{CO} + \text{H}_2\text{O}(\text{l})$), and desorption (step 3: $* \text{CO} \rightarrow * + \text{CO}$). In this approach, the onset potential is numerically equivalent to the additive inverse of the largest positive reaction energy considering steps 1 and 2 only ($U_{\text{onset}} = -\max(\Delta G_1, \Delta G_2)/e^-$) as step 3 is not electrochemical. We note that alternative pathways for CO_2RR to CO in the experimental literature suggest that CO_2 may be activated by an electron transfer prior to its adsorption, and the adsorbed species is stabilized by a hydrated cation close to the surface.^{49–53} Since the modeling of decoupled proton–electron transfers is challenging from a plane-wave DFT standpoint, here we limit ourselves to the standard mechanism⁴⁷ using corrected gas-phase energies.

RESULTS AND DISCUSSION

Pinpointing Errors. The data set used to determine the errors (data set A) consists of 27 molecules involved in the CO_2RR and CORR in which we include at least one

representative molecule of the following functional groups: hydrocarbons, alcohols, carboxylic acids, esters, ethers, aldehydes, and ketones. We included compounds with one to five carbon atoms in the structure (see the full list of compounds in Table S2). Data set A contains the DFT-calculated standard free energy of formation (ΔG_{DFT}^0) of the target molecules (γ) using $\text{C}(\text{s})$, $\text{O}_2(\text{g})$, and $\text{H}_2(\text{g})$ as a reference



For instance, for acetaldehyde, eq 1 is $2\text{C} + \frac{1}{2}\text{O}_2 + 2\text{H}_2 \rightarrow \text{C}_2\text{H}_4\text{O}$. The total errors in the formation energy of each molecule in data set A (ϵ_T) represent the discrepancy between ΔG_{DFT}^0 and ΔG_{exp}^0

$$\epsilon_T = \Delta G_{\text{DFT}}^0 - \Delta G_{\text{exp}}^0 \quad (2)$$

It is worth noting that ϵ_T can either be positive or negative (or zero in case there is a perfect energetic description). As a first approximation, we consider a group additivity-type of scheme⁵⁴ where a given molecule with different functional groups may have different errors present in its ΔG_{DFT}^0 . Thus, the total error (ϵ_T) can be decoupled in the separate contributions of the functional groups present in the molecule (ϵ_i). In mathematical terms, this is expressed as $\epsilon_T \approx \sum_{i=1}^n \epsilon_i$ so that the total error with respect to experiments for a given molecule (ϵ_T) is approximately the sum of the errors inherited from the n functional groups present in the molecule (ϵ_i). As shown in Table S2, data set A is formed by CO, CO_2 , and molecules containing CH_x , hydroxyl, carbonyl, carboxyl, ether, and ester functional groups.

A second data set (data set B) consists of calculated free energies of reaction for the CO_2RR and CORR to produce the molecules in data set A (see Tables S3 and S4). We use data set B to verify whether the corrections implemented in data set A are appropriate. This is the case when there is a decrease in the mean absolute error (MAE) and maximum absolute error (MAX) in data set B as the corrections are successively applied. The free energies of reaction in data set B are grouped in two: first, reactions with CO as a reactant and γ as a product, as shown in eq 3 (see Table S4).



For instance, for acetaldehyde, eq 3 is $2\text{CO} + 3\text{H}_2 \rightarrow \text{C}_2\text{H}_4\text{O} + \text{H}_2\text{O}$. Particular cases are the formation of CO_2 and HCOOH from CO, which follow eq 4



Second, data set B contains reactions with CO_2 as a reactant and γ as a product, as shown in eq 5 (see Table S3).



Equation 5 applied to acetaldehyde is $2\text{CO}_2 + 5\text{H}_2 \rightarrow \text{C}_2\text{H}_4\text{O} + 3\text{H}_2\text{O}$. In these equations, water is considered to be in the gas phase ($\text{H}_2\text{O}(\text{g})$; see Section S3 in the Supporting Information). We categorized the errors for each functional based on organic functional groups ($-\text{CH}_x$, hydroxyl, carbonyl, carboxyl, ether, and ester functional groups) and molecules (in particular, CO and CO_2), as shown in Table 1. For example, acetaldehyde has one $-\text{CH}_x$ ($-\text{CH}_3$) group and one carbonyl ($-\text{CHO}$) group. Table S5 contains the corrections added per exchange–correlation functional and organic functional group.

Table 1. Gas-Phase Error Corrections for the Standard Free Energy of CO₂, CO, and Molecules Containing –C=O– (Carbonyl Groups in Aldehydes and Ketones), –CH_x (Alkanes), and –(C=O)O– (Carboxyl Groups in Carboxylic Acids and Esters) as per xc Functional^a

error	PBE	PW91	RPBE	BEEF-vdW
CO ₂	–0.19	–0.15	–0.46	–0.56
CO	0.24	0.25	–0.07	–0.18
–C=O–	–0.10	–0.10	–0.21	–0.27
–CH _x	0.03	–0.01	0.08	0.21
–(C=O)O–	–0.19	–0.19	–0.27	–0.34 (–0.44)
–OH	–0.04	–0.04	–0.01	–0.14

^aThe two values reported for –(C=O)O– when using BEEF-vdW are for carboxylic acids and esters (the latter in parentheses). All values are in eV.

The errors in the standard free energies (hereafter referred to simply as errors) in Table 1 are xc functional-dependent so that the signs and magnitude change in each case, in line with previous studies.¹⁶ This dependence can be expected because exchange–correlation functionals are fitted for certain applications using different data sets.^{9,55} In the following, we will explain how the errors in Table 1 were determined, taking PBE as an example. Note in passing that the analysis is similar for the other functionals included in this study, and all values are tabulated in Section S4 of the Supporting Information.

To pinpoint the errors, we first determined all deviations (ϵ_T) in the calculated free energies of formation of the molecules in data set A relative to the experimental ones using eq 2. We paid special attention to CO₂ and CO as they are the reactants of CO₂RR and CORR, respectively (all reactions in data set B). For PBE, the error in CO₂ is $\epsilon_T^{\text{CO}_2} = -0.19$ eV, whereas that of CO is $\epsilon_T^{\text{CO}} = 0.24$ eV. Thus, the magnitudes of the two errors are comparable but the signs are opposite. The CO₂ error appears in similar molecules such as HCOOH ($\epsilon_T^{\text{HCOOH}} = -0.19$ eV) and CH₃COOH ($\epsilon_T^{\text{CH}_3\text{COOH}} = -0.15$ eV) and is commonly referred to as the OCO backbone error in the literature.^{18,19,32,34} Previous studies reported corrections of –0.45 eV for RPBE^{19,32} and –0.59 eV for BEEF-vdW,³² which agree well with our values of –0.46 and –0.56 eV, respectively. The small correction of –0.07 eV for CO(g) in RPBE is likely a reflection of RPBE's original fit against CO adsorption energies.³³ We note in passing that simultaneous OCO/H₂ corrections are also available in the literature for BEEF-vdW of 0.33/0.09,³² 0.41/0.09,³⁴ and 0.29/0.10 eV.¹⁸

We continued the correction procedure with the simplest molecules in the list, namely, alkanes (only C–H and single C–C bonds) and observed an increasingly positive error depending on the number of hydrocarbon units (–CH_x) (see Table S6). For PBE, that error is on average $\epsilon_{\text{CH}_x} \approx 0.03$ eV/CH_x. Although small, such an error is cumulative, and therefore, for a molecule with 5 –CH_x units, it becomes $\epsilon_{\text{CH}_x} \approx 0.03 \frac{\text{eV}}{\text{CH}_x} \times 5\text{CH}_x \approx 0.15$ eV. Note that we obtained ϵ_{CH_x} by dividing the error in the formation energy of each alkane by the number of –CH_x units in it and averaging the results for all alkanes in data set A.

Beyond alkanes, one can increase the complexity of the molecules with additional functional groups. For example, we noted that the error for aldehydes and ketones decreased proportionally to the length of the chain. Therefore, to decouple the error associated to carbonyl groups from that of

–CH_x groups, we subtracted from the total error of the molecules the error provided by their –CH_x units (see, for instance, Table S7). In mathematical terms, for a molecule with the formula R₁C = OR₂ (where R₁ and R₂ are either –H or –CH_x units), $\epsilon_T \approx n_C \cdot \epsilon_{\text{CH}_x} + \epsilon_{\text{C=O}}$ where n_C is the number of –CH_x units. To illustrate the use of the formula, consider a total error ($\epsilon_T^{\text{C}_2\text{H}_4\text{O}}$) for acetaldehyde of –0.09 eV and a –CH_x error (ϵ_{CH_x}) of 0.03 eV. The carbonyl-associated error is $\epsilon_{\text{C=O}} = \epsilon_T^{\text{C}_2\text{H}_4\text{O}} - n_C \cdot \epsilon_{\text{CH}_x} = -0.11$ eV. Averaging over all the aldehydes and ketones in this study, we obtained $\epsilon_{\text{C=O}} = -0.10$ eV for PBE.

Table 1 shows the CO and CO₂ errors as well as the average errors determined for the following organic functional groups: –C=O– (aldehydes and ketones), –CH_x (alkanes), –(C=O)O– (carboxylic acids and esters), and –OH (alcohols). Note that the error for –(C=O)O– in PBE is identical to that of CO₂, whereas for PW91, RPBE, and BEEF-vdW, that is not the case as the errors have the same signs but sizably different magnitudes. The error in the –OH group for PBE and PW91 is not large enough to warrant correction for simple alcohols. However, this correction may be needed for polyalcohols and/or in studies focused specifically on methanol and ethanol (see the Supporting Information, Section S4.1.4 for more details).

Before closing this subsection, we stress that a detailed description of the assessment of all errors for every xc functional can be found in Section S4 in the Supporting Information. We note that ethylene, acetylene, ethylene oxide, and dimethyl ether are present in data set A. Since a larger sample of molecules would be necessary to determine the errors corresponding to their respective functional groups (alkenes, alkynes, and (cyclic) ethers), here the corrections for those molecules is limited to the corrections in the reactants only (CO and CO₂).

Implementing Energy Corrections. Data set A was used not only to determine total errors in the formation energies of molecules (ϵ_T) but also to assess the organic group contributions to such errors (ϵ_i). In principle, one can use those errors to correct the formation energies of molecules, the combination of which should lead to accurate reaction energies. In this order of ideas, corrected reaction energies ($\Delta G_{\text{DFT,corr}}^0$) can be calculated as

$$\Delta G_{\text{DFT,corr}}^0 = \Delta G_{\text{DFT}}^0 - \left(\sum \epsilon_T^{\text{P}} - \sum \epsilon_T^{\text{R}} \right) \quad (6)$$

where the sums collect all the errors associated to the reactants (ϵ_T^{R}) and products (ϵ_T^{P}), taking into account the stoichiometric coefficients. For example, consider the reduction of CO₂ to acetic acid: $2\text{CO}_2 + 4\text{H}_2 \rightarrow \text{CH}_3\text{COOH} + 2\text{H}_2\text{O}$. We find with RPBE that $\Delta G_{\text{DFT}}^0 = 0.32$ eV, whereas $\Delta G_{\text{exp}}^0 = -0.44$ eV, which corresponds to a large total error of $\epsilon_T^{\text{CH}_3\text{COOH}} = 0.76$ eV. According to Table 1, RPBE has errors associated to the description of CO₂, the –COOH group, and the –CH₃ moiety in CH₃COOH. If the errors pinpointed using data set A are indeed contributing to the large total error, then suitably correcting CO₂ and CH₃COOH should lead to a sizable reduction of the total error. This is what we find as $\sum \epsilon_T^{\text{R}} = 2\epsilon_T^{\text{CO}_2} = -0.92$ eV and $\sum \epsilon_T^{\text{P}} = \epsilon_{\text{CH}_3} + \epsilon_{\text{C=O}} = -0.19$ eV so that $\Delta G_{\text{DFT,corr}}^0 = -0.41$ eV, which differs from the experimental value ($\Delta G_{\text{exp}}^0 = -0.44$ eV) by 0.03 eV only.

To verify that the errors in the reaction energies of data set B are systematically reduced upon applying the corrections in Table 1, we followed a stepwise procedure. First, we applied

corrections to data set B only related to reactants (namely, CO₂ and CO). Next, we applied corrections related to products. Figures 1 and 2 show the calculated free energies of

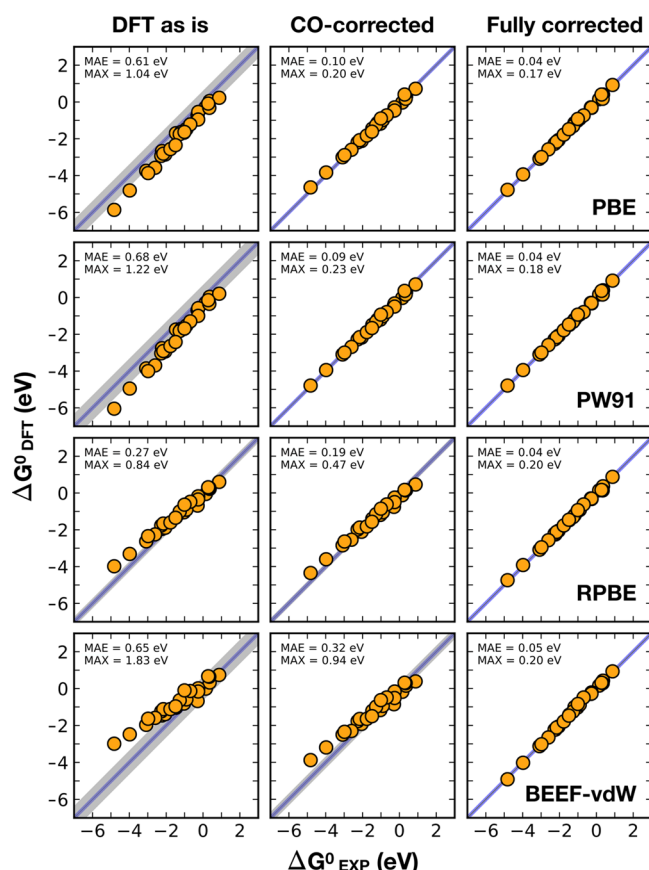


Figure 1. Parity plots for the experimental and DFT-calculated free energies of production of 27 different compounds from CO and H₂ using PBE, PW91, RPBE, and BEEF-vdW. The left column shows the data calculated with DFT without any correction. The center column shows the data upon the first correction (errors in CO), and the right column shows the data after correcting for errors in CO and the products. The mean and maximum absolute errors (MAE and MAX) are shown in each case. The shaded gray area is \pm MAE in each case. The blue shaded area around the parity line covers an area of ± 0.15 eV.

reaction versus the experimental free energies for the four functionals studied (PBE, PW91, RPBE, and BEEF-vdW). Figure 1 provides parity plots for CO-based reactions (eqs 3 and 4), and Figure 2 does so for CO₂-based reactions (eq 5). From the three columns in each figure, the first one corresponds to the noncorrected DFT data, the plots in the second column contain the data upon correcting for reactant-related errors (namely, CO or CO₂), and the third column contains the data upon correcting for reactant- and product-related errors altogether.

More molecules can be added to data set A so as to include more organic functional groups and molecules with several groups in their structure. Molecules with alkene, alkyne, epoxy, and ether functional groups as well as aromatic compounds are necessary in data set A to determine their corresponding errors. Here, the free energies of production from CO or CO₂ of ethylene, acetylene, dimethyl ether, and ethylene oxide were corrected for the errors in the reactants only, and no product-related corrections were made (see Table S5).

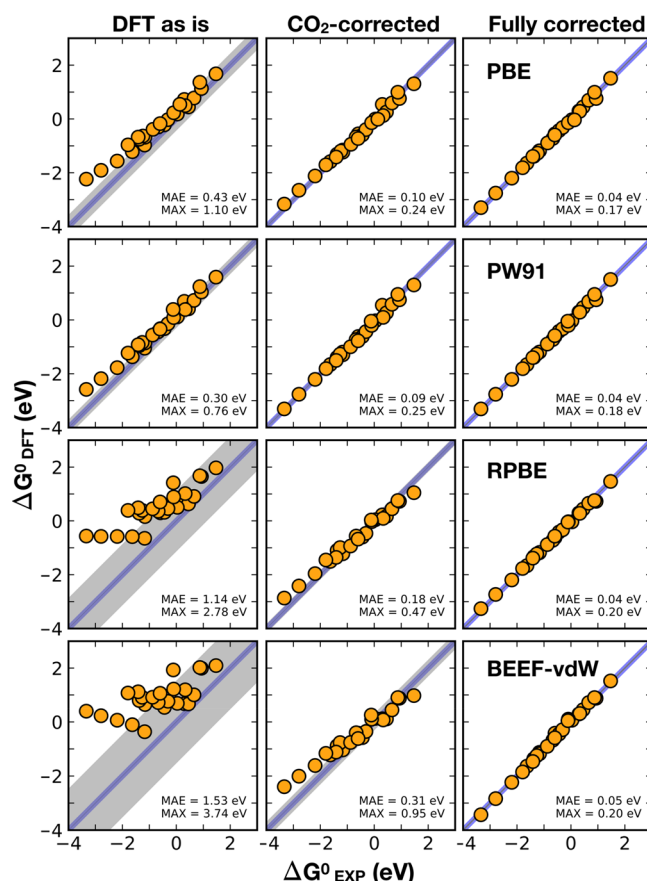


Figure 2. Parity plot for the experimental and DFT-calculated free energies of production of 27 different products from CO₂ and H₂ using PBE, PW91, RPBE, and BEEF-vdW. The left column shows the data calculated with DFT without any correction. The center column shows the data upon the first correction (errors in CO₂), and the right column shows the data after correcting for errors in CO₂ and the products. The mean and maximum absolute errors (MAE and MAX) are shown in each case. The shaded gray area is \pm MAE in each case. The blue shaded area around the parity line covers an area of ± 0.15 eV.

The gray-shaded areas in Figures 1 and 2 cover an area around the parity line of \pm MAE, and the purple-shaded area extends over ± 0.15 eV around the parity line. For CO reduction reactions and PBE calculations, the MAE is initially 0.61 eV (left column) and is lowered to 0.10 eV after applying the CO correction (central column) and to 0.04 eV after applying both CO and product-related corrections (right column). Similarly, the MAXs go from 1.04 to 0.20 and then to 0.17 eV. For the CO₂ reduction reactions and PBE, the MAE is successively reduced from 0.43 to 0.10 and then to 0.04 eV. Likewise, the MAXs decrease from 1.10 to 0.24 and finally to 0.17 eV. Further details can be found in Table S24 where the MAEs after the first and second correction for all the xc functionals are provided. We conclude from those values that the errors in data set B are lowered by 1 order of magnitude once the correction scheme is applied to the species in data set A.

An alternative analysis splitting data set A into a training set and an extrapolation set can be found in Section S7 in the Supporting Information. We find approximately the same functional-related errors as in Table 1 (within ± 0.01 eV on average). The MAEs in the extrapolation set after the

corrections are comparable to those in Figures 1 and 2, illustrating the predictive power of the method and its statistical reliability.

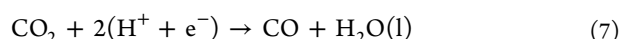
Applications in Electrocatalysis. Table 2 reveals an important commonality among the xc functionals under study:

Table 2. CO₂ and CO Errors and their Nearly Constant Difference ($\epsilon_T^{\text{CO}} - \epsilon_T^{\text{CO}_2}$) across xc Functionals^a

error	PBE	PW91	RPBE	BEEF-vdW
CO	0.24	0.25	−0.07	−0.18
CO ₂	−0.19	−0.15	−0.46	−0.56
$\epsilon_T^{\text{CO}} - \epsilon_T^{\text{CO}_2}$	0.43	0.40	0.39	0.38
average	0.40			
standard deviation	0.02			

^aAll values are in eV.

although the CO and CO₂ errors change from one functional to the next, their difference is nearly constant and equal to ~0.4 eV, on average. This constant energetic separation poses a fundamental limitation for the modeling of catalytic reactions wherein those two compounds are involved, one as a reactant and the other as a product. To show the reaches of this finding, let us consider the example of CO₂ electrocatalytic reduction (CO₂RR) to CO



The backwards reaction is known as CO oxidation and is also an important electrocatalytic reaction involved in direct ethanol and methanol fuel cells.⁵⁶ Moreover, eq 7 can also be catalyzed in the gas phase using H₂ in a process called reverse water–gas shift, and the backwards reaction is the industrial process known as the water–gas shift.⁵⁷ In brief, DFT-based models of this seemingly simple process with numerous applications in electrocatalysis and heterogeneous catalysis may have large gas-phase associated errors.

Indeed, Figure 3 compares CO₂RR to CO on Au(111) single-crystal electrodes using PBE with (Figure 3b) and without (Figure 3a) gas-phase corrections applied to CO₂ and CO. Likewise, Figures S5 and S6 in the Supporting Information, Section S5 provide the data for Au(100) and Au(110). In Figure 3a, where DFT data appear as is, the reaction energy of eq 7 is 0.63 eV. Conversely, it is 0.20 eV in

Figure 3b where the energies of CO₂ and CO have been corrected. For comparison, such difference is 0.20 eV in experiments²⁵ (it is 0.30 eV in Table S3). The difference stems from the liquid state of water in eq 7). In terms of the equilibrium potential of the reaction, this all means that PBE predicts it to be at −0.32 V versus RHE, whereas both the correction method and experiments set it at −0.10 V versus RHE. The difference is substantial and amounts to ~220 mV. Note in passing that there are no changes in the energy differences between *COOH and *CO as the corrections are only applied to the gas phase. Although corrections for adsorbates have been proposed before,¹⁸ they escape the subject and scope of this article.

Within the context of CO₂RR modeling with the computational hydrogen electrode,^{47,48} the onset potential is given by the largest positive consecutive difference in Figure 3 ($U_{\text{onset}} = -\max(\Delta G_1, \Delta G_2)/e^-$; see the Computational Methods section). In Figure 3a, such a difference is 0.90 eV, whereas in Figure 3b, it is 0.71 eV so that the predicted onset potentials are −0.90 and −0.71 V versus RHE, respectively. As the experimental value of the onset potential is −0.66 V versus RHE,⁵⁸ the deviations from experiments are ~0.24 (as is) and 0.05 V (corrected).

We note that the sizable lowering of the error from 0.24 to 0.05 V is a direct result of correcting gas-phase energetics. To assess whether this is a particularity of Au(111) electrodes or part of a more general trend, we also compared the calculated and experimental onset potentials for Au(100), Au(110), Au_{poly}, Ag(111), Ag_{poly}, and Cu_{poly}. The results in Figure 4a show that DFT data are systematically deviated from the parity line, which results in a MAE of 0.20 V and a MAX of 0.27 V. Conversely, the CO₂ and CO corrected data in Figure 4b are located around the parity line with MAE = 0.06 V and MAX = 0.09 V. Substantial improvements are also observed for Au(111) and Au(100) using gas-phase corrections with RPBE (see Figures S7 and S8). Thus, we conclude that models for CO₂RR to CO may in general benefit from the gas-phase corrections found in this work.

CONCLUSIONS

When interfaces between metals and fluids are simulated at the GGA level, sizable errors may appear in the description of the gas-phase molecules. Here, we proposed a two-step semi-

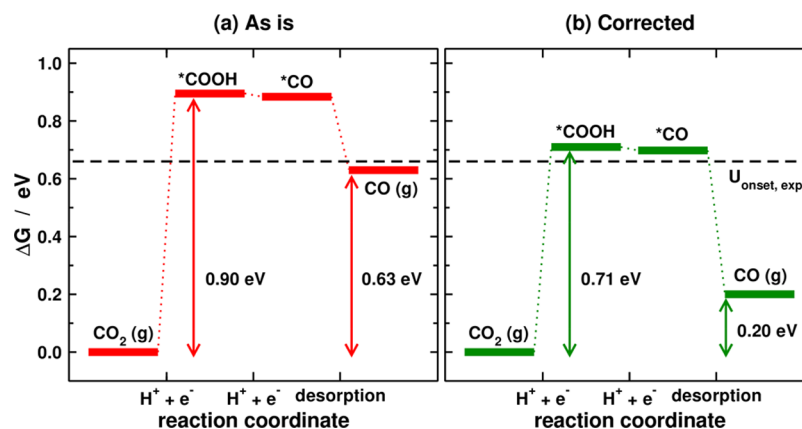


Figure 3. Free energy diagrams for CO₂ reduction to CO using Au(111) single-crystal electrodes. (a) Using DFT-PBE data as is and (b) correcting CO₂ and CO for their gas-phase errors. The black dashed line at 0.66 eV marks the free energy corresponding to the experimental onset potential of −0.66 V vs RHE.⁵⁸

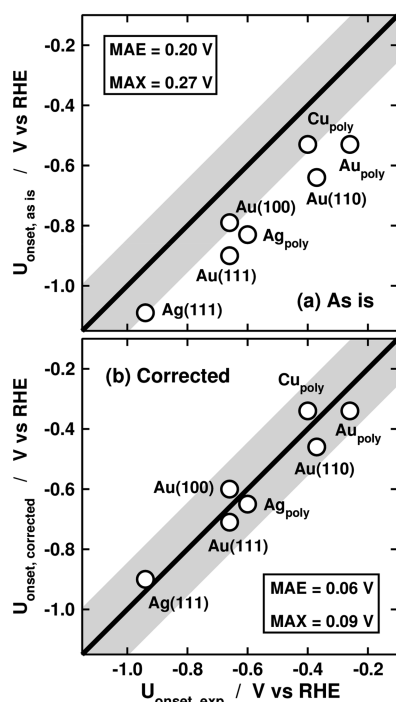


Figure 4. Parity plots comparing the onset potentials for electrochemical CO_2 reduction to CO using different metals. (a) Using DFT-PBE data as is and (b) correcting CO_2 and CO for their gas-phase errors. The mean and maximum absolute errors (MAE and MAX) are shown in each case. The gray areas around the parity line cover an area of ± 0.15 V around it. The experimental data were taken from refs 58–61 (see details in Table S25).

empirical method to determine gas-phase errors based on the formation energies of 27 different molecules. Furthermore, implementing the corresponding corrections allow for predictions in the analyzed data set of CO_2RR and CORR reaction energies that lower by 1 order of magnitude the average and maximum errors with respect to experiments.

The method also shows that the errors for CO_2 and CO differ by ~ 0.4 eV for all the examined exchange–correlation functionals. Thus, an intrinsic limitation of DFT exists for the accurate description of reaction energies containing these two molecules, as is the case for CO_2 reduction to CO, among others. Such limited description leads to inaccurate predictions of equilibrium and onset potentials, which may hinder the rational design of catalysts.

Conversely, using our correction scheme on various Au, Ag, and Cu electrodes decreased the average error in the predicted onset potentials from 0.20 to 0.06 V with respect to experiments. Therefore, in addition to pinpointing and lowering gas-phase errors, the method also helps in providing more accurate electrocatalytic models.

While the present corrections have been applied for electrochemical reactions, the procedure is general enough to be applied to correct the thermochemistry of heterogeneously catalyzed reactions where reactants and products are in the gas phase but the overall reaction takes place at the catalyst's surface. Finally, the correction protocol can be enriched by adding more gas-phase molecules to the data set and using machine learning algorithms to detect and predict errors in structurally more complex substances.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acscatal.0c01075>.

Plane-wave convergence test, tabulated data for data sets A and B, how to pinpoint various errors for PBE, PW91, RPBE, and BEEF-vdW exchange–correlation functionals, additional electrocatalysis-related figures and tabulated data, and alternative error analysis using training and extrapolation sets (PDF)

■ AUTHOR INFORMATION

Corresponding Authors

Marc T. M. Koper – Leiden Institute of Chemistry, Leiden University, Leiden 2300RA, The Netherlands; orcid.org/0000-0001-6777-4594; Email: m.koper@lic.leidenuniv.nl

Federico Calle-Vallejo – Departament de Ciència de Materials i Química Física & Institut de Química Teòrica i Computacional (IQTUB), Universitat de Barcelona, Barcelona 08028, Spain; orcid.org/0000-0001-5147-8635; Email: f.calle.vallejo@ub.edu

Authors

Laura P. Granda-Marulanda – Leiden Institute of Chemistry, Leiden University, Leiden 2300RA, The Netherlands; Departament de Ciència de Materials i Química Física & Institut de Química Teòrica i Computacional (IQTUB), Universitat de Barcelona, Barcelona 08028, Spain

Alejandra Rendón-Calle – Departament de Ciència de Materials i Química Física & Institut de Química Teòrica i Computacional (IQTUB), Universitat de Barcelona, Barcelona 08028, Spain; Departamento de Ingeniería de Procesos, Universidad EAFIT, 050022, Colombia

Santiago Builes – Departamento de Ingeniería de Procesos, Universidad EAFIT, 050022, Colombia; orcid.org/0000-0003-4273-3774

Francesc Illas – Departament de Ciència de Materials i Química Física & Institut de Química Teòrica i Computacional (IQTUB), Universitat de Barcelona, Barcelona 08028, Spain; orcid.org/0000-0003-2104-6123

Complete contact information is available at: <https://pubs.acs.org/doi/10.1021/acscatal.0c01075>

Notes

The authors declare no competing financial interest.

■ ACKNOWLEDGMENTS

This work was performed under the project HPC-EUROPA3 (INFRAIA-2016-1-730897) with the support of the EC Research Innovation Action under the H2020 program and the computer resources and technical support provided by the Barcelona Supercomputing Center. We acknowledge the financial support from Spanish MICIUN through RTI2018-095460-B-I00 and María de Maeztu MDM-2017-0767 grants and, in part, from Generalitat de Catalunya, grant 2017SGR13. F.C.-V. thanks MICIUN for a Ramón y Cajal research contract (RYC-2015-18996), and F.I. acknowledges additional support from the 2015 ICREA Academia Award for Excellence in University Research. L.P.G.-M. and M.T.M.K. acknowledge funding from the European Union through the A-leaf project (732840-A-LEAF). This work was also supported by Universidad EAFIT through project 690-000048. The use of

supercomputing facilities at SURFsara was sponsored by NWO Physical Sciences with financial support by NWO. We also used supercomputing resources of the Centro de Computación Científica Apolo at Universidad EAFIT (<http://www.eafit.edu.co/apolo>), the Center for Functional Nanomaterials, a U.S. DOE Office of Science Facility, the Scientific Data and Computing Center, a component of the Computational Science Initiative, at Brookhaven National Laboratory under contract no. DE-SC0012704, and Red Española de Supercomputación (RES) at SCAYLE (projects QS-2019-3-0018, QS-2019-2-0023, and QCM-2019-1-0034).

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