A Semiempirical Method to Detect and Correct DFT-Based Gas-Phase Errors and Its Application in Electrocatalysis

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ABSTRACT: Computational models of adsorption at metal surfaces are often based on DFT and make use of the generalized gradient approximation. This likely implies the presence of sizable errors in the gas-phase energetics. Here, we take a step closer toward chemical accuracy with a semiempirical method to correct the gas-phase energetics of PBE, PW91, RPBE, and BEEF-vDw exchange–correlation functionals. The proposed two-step method is tested on a data set of 27 gas-phase molecules belonging to the carbon cycle: first, the errors are pinpointed based on formation energies, and second, the respective corrections are sequentially applied to ensure the progressive lowering of the data set’s mean and maximum errors. We illustrate the benefits of the method in electrocatalysis by a substantial improvement of the calculated equilibrium and onset potentials for CO₂ reduction to CO on Au, Ag, and Cu electrodes. This suggests that fast and systematic gas-phase corrections can be devised to augment the predictive power of computational catalysis models.

KEYWORDS: gas-phase errors, gas-phase corrections, carbon monoxide, carbon dioxide, DFT, electrochemical reduction of CO₂

INTRODUCTION

For decades, considerable effort has been devoted to increasing the accuracy of density functional theory (DFT). This has been done by developing more accurate exchange–correlation functionals at the generalized gradient approximation (GGA) level, hybrid functionals, and range-separated functionals. In addition, different correction schemes have been developed to account for electron localization or dispersion interactions. Lately, machine learning schemes have also been proposed to bypass Kohn–Sham equations. In general, these efforts include careful computational benchmarking and comparison to experiments.

An agreement has been reached in the scientific community about the level of theory required to simulate certain materials with a good tradeoff between computational time and accuracy. For instance, hybrid functionals are advisable for molecules and solids with localized electrons, while GGAs usually suffice for bulk and surface metals. However, the choice is not trivial when dealing with systems where metals and molecules are involved and ought to be simulated at the same level of theory. In such a case, the accuracy may be improved by using GGA functionals and adding semiempirical corrections to the DFT energies of molecules, as done for thermochemical reaction energies of interest in catalysis, formation and decomposition energies of solids, and catalytic kinetic barriers.

In this article, we provide a simple and fast procedure for detecting gas-phase errors based on the formation energies of reactants and products calculated with DFT. Improving the description of the gas phase is shown to enhance catalytic predictive power by analyzing the electrocatalytic CO₂ reduction reaction to CO on Au, Ag, and Cu electrodes. The reduction of CO₂ and CO (hereafter denoted as CO₂RR and CORR, respectively) are of great importance in catalysis science and technology as they lead to valuable feedstocks and fuels such as methane, ethylene, ethanol, and formic acid while helping in balancing the carbon cycle.

Although DFT has been used to predict enhanced catalysts for other electrocatalytic reactions, it has been, so far, challenging to elaborate robust design routines for CO₂RR and CORR to hydrocarbons and oxygenates. Thus, the method presented here may help boost materials design via screening for those paramount reactions.

COMPUTATIONAL METHODS

All calculations were performed using the Vienna Ab initio simulation package. Dissimilar gas-phase errors have been pointed out in previous studies for the total energy of CO(g) and CO₂(g) using PBE and RPBE. In addition, others...
suggested a correction for the total energy of H₂(g) to be applied only when using BEEF-vdW. Thus, we made a functional-dependent analysis including four different xc functionals habitually used in catalysis, namely, PBE, PW91, RPBE, and BEEF-vdW. The gas-phase molecules were relaxed with the conjugate gradient algorithm in boxes of ~3375 Å³, considering only the Γ point. The effect of the cores on the valence electron density is incorporated using the projector-augmented wave (PAW) method. To compute the formation energies of the molecules, graphite was represented by graphene. Approximating graphene as the standard state of carbon is based on the weak interlayer cohesive energy of graphite (0.031–0.064 eV/atom) (see Section S6 in the Supporting Information). The optimized interatomic distances of graphene are 1.43 Å (PBE and RPBE) and 1.42 Å (PW91 and BEEF-vdW).

The convergence criterion for the maximal forces on the atoms for all simulations was 0.01 eV Å⁻¹, and the plane-wave cutoff was set to 400 eV. Convergence tests for the free energy of reaction of CO₂(g) + H₂(g) → CO(g) + H₂O(g) with plane-wave cutoffs in the range of 300–1000 eV within PBE showed that 400 eV is enough to achieve accurate reaction energies with an average difference of ~5 meV (see Table S1). None of the species analyzed had unpaired electrons, so spin-unrestricted calculations were not required. Gaussian smearing with k_BT = 0.001 eV was used. In all cases, the energies were extrapolated to 0 K.

The reaction free energies were obtained as ΔG° = ΔG_DFT + ΔZPE − TΔS° where ZPE is the zero-point energy contribution calculated from the vibrational frequencies obtained using the harmonic oscillator approximation. The standard total entropies (S°) and the experimental standard free energies (ΔG°_exp) were obtained from thermodynamic tables at T = 298.15 K. In cases where ΔG°_exp was not tabulated, it was evaluated by combining entropy and enthalpy values: ΔG°_exp = ΔH°_exp − TΔS°_exp. We did not include heat capacity effects as recent studies showed that formation energies are not significantly modified by them from 0 to 298.15 K.

Electrocatalytic CO₂ reduction to CO was modeled based on the free energy scheme described in previous reports, making use of the computational hydrogen electrode for the description of proton–electron transfers. The reaction pathway proceeds via CO₂ hydrogenation (step 1: CO₂ + H⁺ + e⁻ → *COOH), followed by *CO formation (step 2: *COOH + H⁺ + e⁻ → *CO + H₂O()), and desorption (step 3: *CO → CO). In this approach, the onset potential is numerically equivalent to the additive inverse of the largest positive reaction energy considering steps 1 and 2 only (U_onset = −max(ΔG_i, ΔG_j)/e⁻) as step 3 is not electrochemical. We note that alternative pathways for CO₂RR to CO in the experimental literature suggest that CO₂ may be activated by an electron transfer prior to its adsorption, and the adsorbed species is stabilized by a hydrated cation close to the surface. Since the modeling of decoupled proton–electron transfers is challenging from a plane-wave DFT standpoint, here we limit ourselves to the standard mechanism using corrected gas-phase energies.

### RESULTS AND DISCUSSION

#### Pinpointing Errors

The data set used to determine the errors (data set A) consists of 27 molecules involved in the CO₂RR and CORR in which we include at least one representative molecule of the following functional groups: hydrocarbons, alcohols, carboxylic acids, esters, ethers, aldehydes, and ketones. We included compounds with one to five carbon atoms in the structure (see the full list of compounds in Table S2). Data set A contains the DFT-calculated standard free energy of formation (ΔG_DFT) of the target molecules (γ) using C(s), O₂(g), and H₂(g) as a reference

\[ \text{aC + bO}_2 + cH_2 \rightarrow \gamma \]  \hspace{1cm} (1)

For instance, for acetaldehyde, eq 1 is \( 2\text{C} + \frac{1}{2}\text{O}_2 + 2\text{H}_2 \rightarrow \text{C}_2\text{H}_4\text{O} \). The total errors in the formation energy of each molecule in data set A (ε_T) represent the discrepancy between ΔG_DFT and ΔG_exp

\[ \varepsilon_T = \Delta G_DFT - \Delta G_{exp} \]  \hspace{1cm} (2)

It is worth noting that ε_T can either be positive or negative (or zero in case there is a perfect energetic description). As a first approximation, we consider a group additivity-type of scheme where a given molecule with different functional groups may have different errors present in its ΔG_DFT. Thus, the total error (ε_T) can be decoupled in the separate contributions of the functional groups present in the molecule (ε_i). In mathematical terms, this is expressed as ε_T ≈ ∑n_i=1 ε_i so that the total error with respect to experiments for a given molecule (ε_T) is approximately the sum of the errors inherited from the n functional groups present in the molecule (ε_i). As shown in Table S2, data set A is formed by CO, CO₂, and molecules containing CH₃, hydroxyl, carboxyl, carbonyl, ether, and ester functional groups.

A second data set (data set B) consists of calculated free energies of reaction for the CO₂RR and CORR to produce the molecules in data set A (see Tables S3 and S4). We use data set B to verify whether the corrections implemented in data set A are appropriate. This is the case when there is a decrease in the mean absolute error (MAE) and maximum absolute error (MAX) in data set B as the corrections are successively applied. The free energies of reaction in data set B are grouped in two: first, reactions with CO as a reactant and γ as a product, as shown in eq 3 (see Table S4).

\[ \text{gCO + kH}_2 \rightarrow \gamma + m\text{H}_2\text{O} \]  \hspace{1cm} (3)

For instance, for acetaldehyde, eq 3 is 2CO + 3H₂ → C₂H₄O + H₂O. Particular cases are the formation of CO₂ and HCOOH from CO, which follow eq 4

\[ \text{CO + H}_2\text{O} \rightarrow \gamma + r\text{H}_2 \]  \hspace{1cm} (4)

Second, data set B contains reactions with CO₂ as a reactant and γ as a product, as shown in eq 5 (see Table S3).

\[ x\text{CO}_2 + y\text{H}_2 \rightarrow \gamma + z\text{H}_2\text{O} \]  \hspace{1cm} (5)

Equation 5 applied to acetaldehyde is 2CO₂ + 5H₂ → C₂H₄O + 3H₂O. In these equations, water is considered to be in the gas phase (H₂O(g); see Section S3 in the Supporting Information). We categorized the errors for each functional based on organic functional groups (−CH₂ hydroxyl, carbonyl, carbonyl, ether, and ester functional groups) and molecules (in particular, CO and CO₂), as shown in Table 1. For example, acetaldehyde has one −CH₂ (−CH₃) group and one carbonyl (−CHO) group. Table S5 contains the corrections added per exchange–correlation functional and organic functional group.
The errors in the standard free energies (hereafter referred to simply as errors) in Table 1 are xc functional-dependent so that the signs and magnitude change in each case, in line with previous studies. This dependency can be expected because exchange–correlation functionals are fitted for certain applications using different data sets. In the following, we will explain how the errors in Table 1 were determined, taking PBE as an example. Note in passing that the analysis is similar for the other functionals included in this study, and all values are tabulated in Section S4 of the Supporting Information.

To pinpoint the errors, we first determined all deviations ($\epsilon_T$) in the calculated free energies of formation of the molecules in data set A relative to the experimental ones using eq 2. We paid special attention to CO$_2$ and CO as they are the reactants of CO$_2$RR and CORR, respectively (all reactions in data set B). For PBE, the error in CO$_2$ is $\epsilon_{T}^{CO_2} = -0.19$ eV, whereas that of CO is $\epsilon_{T}^{CO} = 0.24$ eV. Thus, the magnitudes of the two errors are comparable but the signs are opposite. The CO$_2$ error appears in similar molecules such as HCOOH ($\epsilon_{T}^{HCOOH} = -0.19$ eV) and CH$_2$COOH ($\epsilon_{T}^{CH_2COOH} = 0.15$ eV) and is commonly referred to as the OCO backbone error in the literature. Previous studies reported corrections of $-$0.45 eV for RPBE $^{10}$, $-$0.59 eV for BEEF-vdW, $^{32}$ which agree well with our values of $-$0.46 and $-$0.56 eV, respectively. The small correction of $-$0.07 eV for CO(g) in RPBE is likely a reflection of RPBE’s original fit against CO adsorption energies. We note in passing that simultaneous OCO/H$_2$ corrections are also available in the literature for BEEF-vdW of 0.33/0.09, $^{32}$ 0.41/0.09, $^{34}$ and 0.29/0.10 eV. $^{18}$

We continued the correction procedure with the simplest molecules in the list, namely, alkanes (only C–H and single C–C bonds) and observed an increasingly positive error depending on the number of hydrocarbon units (–CH$_2$) (see Table S6). For PBE, the error is on average $\epsilon_{T}^{alkane} \approx 0.03$ eV/CH$_2$. Although small, such an error is cumulative, and therefore, for a molecule with 5 –CH$_2$ units, it becomes $\epsilon_{T}^{alkane} \approx 0.02$ eV/CH$_2 = 0.15$ eV. Note that we obtained $\epsilon_{T}^{alkane}$ by dividing the error in the formation energy of each alkane by the number of –CH$_2$ units in it and averaging the results for all alkanes in data set A.

Beyond alkanes, one can increase the complexity of the molecules with additional functional groups. For example, we noted that the error for aldehydes and ketones decreased proportionally to the length of the chain. Therefore, to decouple the error associated to carbonyl groups from that of –CH$_2$ groups, we subtracted from the total error of the molecules the error provided by their –CH$_{n}$ units (see, for instance, Table S7). In mathematical terms, for a molecule with the formula R$_i$C = OR$_j$ (where R$_i$ and R$_j$ are either –H or –CH$_{n}$ units), $\epsilon_{T} = n_{C} \epsilon_{T}^{CH_{n}} + \epsilon_{T}^{C=O} - \epsilon_{T}^{C=O} - \epsilon_{T}^{CO}$ (where $n_{C}$ is the number of –CH$_{n}$ units. To illustrate the use of the formula, consider a total error ($\epsilon_{T}^{CH_{n}}$) for acetaldehyde of $-$0.09 eV and a –CH$_{3}$ error ($\epsilon_{T}^{CH_{3}}$) of 0.03 eV. The carbonyl-associated error is $\epsilon_{T}^{C=O} = \epsilon_{T}^{CH_{3}} - 0.09 eV$ for acetaldehyde and $\epsilon_{T}^{C=O} = -0.11$ eV. Averaging over all the aldehydes and ketones in this study, we obtained $\epsilon_{T}^{C=O} = -0.10$ eV for PBE.

Table 1 shows the CO and CO$_2$ errors as well as the average errors determined for the following organic functional groups: –C=O– (aldehydes and ketones), –CH$_{n}$ (alkanes), –(C=O)O– (carboxylic acids and esters), and –OH (alcohols). Note that the error for $-(C=O)O-$ in PBE is identical to that of CO$_2$, whereas for PW91, RPBE, and BEEK-vdW, that is not the case as the errors have the same signs but sizably different magnitudes. The error in the –OH group for PBE and PW91 is not large enough to warrant correction for simple alcohols. However, this correction may be needed for polyalcohols and/or in studies focused specifically on methanol and ethanol (see the Supporting Information, Section S4.1.4 for more details).

Before closing this subsection, we stress that a detailed description of the assessment of all errors for every xc functional can be found in Section S4 in the Supporting Information. We note that ethylene, acetylene, ethylene oxide, and dimethyl ether are present in data set A. Since a larger sample of molecules would be necessary to determine the errors corresponding to their respective functional groups (alkenes, alkynes, and (cyclic) ethers), here the corrections for those molecules is limited to the corrections in the reactants only (CO and CO$_2$).

### Implementing Energy Corrections

Data set A was used not only to determine total errors in the formation energies of molecules ($\epsilon_{T}^{p}$) but also to assess the organic group contributions to such errors ($\epsilon_{i}^{p}$). In principle, one can use those errors to correct the formation energies of molecules, the combination of which should lead to accurate reaction energies. In this order of ideas, corrected reaction energies ($\Delta G_{DFT,corr}$) can be calculated as

$$\Delta G_{DFT,corr} = \Delta G_{DFT} - \left( \sum \epsilon_{T}^{p} - \sum \epsilon_{T}^{p} \right)$$

(6)

where the sums collect all the errors associated to the reactants ($\epsilon_{T}^{p}$) and products ($\epsilon_{T}^{p}$), taking into account the stoichiometric coefficients. For example, consider the reduction of CO$_2$ to acetic acid: 2CO$_2$ + 4H$_2$ → CH$_3$COOH + 2H$_2$O. We find with RPBE that $\Delta G_{DFT} = 0.32$ eV, whereas $\Delta G_{esp} = -0.44$ eV, which corresponds to a large total error of $\Delta G_{esp} = 0.76$ eV. According to Table 1, RPBE has errors associated to the description of CO$_2$, the –COOH group, and the –CH$_3$ moiety in CH$_3$COOH. If the errors pinpointed using data set A are indeed contributing to the large total error, then suitably correcting CO$_2$ and CH$_3$COOH should lead to a sizable reduction of the total error. This is what we find as $\sum \epsilon_{T}^{p} = -2 \epsilon_{T}^{CO_2} = -0.92$ eV and $\sum \epsilon_{T}^{p} = \epsilon_{T}^{CH_3} + \epsilon_{T}^{C=O} - \epsilon_{T}^{O} = -0.19$ eV so that $\Delta G_{DFT,corr} = -0.41$ eV, which differs from the experimental value ($\Delta G_{esp} = -0.44$ eV) by 0.03 eV only.

To verify that the errors in the reaction energies of data set B are systematically reduced upon applying the corrections in Table 1, we followed a stepwise procedure. First, we applied
corrections to data set B only related to reactants (namely, CO and CO). Next, we applied corrections related to products. Figures 1 and 2 show the calculated free energies of reaction versus the experimental free energies for the four functionals studied (PBE, PW91, RPBE, and BEEF-vdW).

Figure 1 provides parity plots for CO-based reactions (eqs 3 and 4), and Figure 2 does so for CO₂-based reactions (eq 5). From the three columns in each figure, the first one corresponds to the noncorrected DFT data, the plots in the second column contain the data upon correcting for reactant-related errors (namely, CO or CO₂), and the third column contains the data upon correcting for reactant- and product-related errors altogether.

More molecules can be added to data set A so as to include more organic functional groups and molecules with several groups in their structure. Molecules with alkene, alkyne, epoxy, and ether functional groups as well as aromatic compounds are necessary in data set A to determine their corresponding errors. Here, the free energies of production from CO or CO₂ of ethylene, acetylene, dimethyl ether, and ethylene oxide were corrected for the errors in the reactants only, and no product-related corrections were made (see Table S5).

The gray-shaded areas in Figures 1 and 2 cover an area around the parity line of ± MAE, and the purple-shaded area extends over ±0.15 eV around the parity line. For CO reduction reactions and PBE calculations, the MAE is initially 0.61 eV (left column) and is lowered to 0.10 eV after applying the CO correction (central column) and to 0.04 eV after applying both CO and product-related corrections (right column). Similarly, the MAXs go from 1.04 to 0.20 and then to 0.17 eV. For the CO₂ reduction reactions and PBE, the MAE is successively reduced from 0.43 to 0.10 and then to 0.04 eV. Likewise, the MAXs decrease from 1.10 to 0.24 and finally to 0.17 eV. Further details can be found in Table S24 where the MAEs after the first and second correction for all the xc functionals are provided. We conclude from those values that the errors in data set B are lowered by 1 order of magnitude once the correction scheme is applied to the species in data set A.

An alternative analysis splitting data set A into a training set and an extrapolation set can be found in Section S7 in the Supporting Information. We find approximately the same functional-related errors as in Table 1 (within ±0.01 eV on average). The MAEs in the extrapolation set after the
corrections are comparable to those in Figures 1 and 2, illustrating the predictive power of the method and its statistical reliability.

**Applications in Electrocatalysis.** Table 2 reveals an important commonality among the xc functionals under study:

Table 2. CO$_2$ and CO Errors and their Nearly Constant Difference ($\varepsilon_f^{\text{CO}} - \varepsilon_f^{\text{CO}_2}$) across xc Functionals.$^{44}$

<table>
<thead>
<tr>
<th>Error</th>
<th>PBE</th>
<th>PW91</th>
<th>RPBE</th>
<th>BEEF-vdW</th>
</tr>
</thead>
<tbody>
<tr>
<td>CO</td>
<td>0.24</td>
<td>0.25</td>
<td>−0.07</td>
<td>−0.18</td>
</tr>
<tr>
<td>CO$_2$</td>
<td>−0.43</td>
<td>−0.15</td>
<td>−0.46</td>
<td>−0.56</td>
</tr>
<tr>
<td>$\varepsilon_f^{\text{CO}} - \varepsilon_f^{\text{CO}_2}$</td>
<td>0.43</td>
<td>0.40</td>
<td>0.39</td>
<td>0.38</td>
</tr>
<tr>
<td>Average</td>
<td>0.40</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Standard deviation</td>
<td>0.02</td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

“All values are in eV.

although the CO and CO$_2$ errors change from one functional to the next, their difference is nearly constant and equal to ~0.4 eV, on average. This constant energetic separation poses a fundamental limitation for the modeling of catalytic reactions wherein those two compounds are involved, one as a reactant and the other as a product. To show the reaches of this finding, let us consider the example of CO$_2$ electrocatalytic reduction (CO$_2$RR) to CO

$$\text{CO}_2 + 2(\text{H}^+ + \text{e}^-) \rightarrow \text{CO} + \text{H}_2\text{O}(l)$$ (7)

The backwards reaction is known as CO oxidation and is also an important electrocatalytic reaction involved in direct ethanol and methanol fuel cells. Moreover, eq 7 can also be catalyzed in the gas phase using H$_2$ in a process called reverse water–gas shift, and the backwards reaction is the industrial process known as the water–gas shift. In brief, DFT-based models of this seemingly simple process with numerous applications in electrocatalysis and heterogeneous catalysis may have large gas-phase associated errors.

Indeed, Figure 3 compares CO$_2$RR to CO on Au(111) single-crystal electrodes using PBE with (Figure 3b) and without (Figure 3a) gas-phase corrections applied to CO$_2$ and CO. Likewise, Figures S5 and S6 in the Supporting Information, Section S5 provide the data for Au(100) and Au(110). In Figure 3a, where DFT data appear as is, the reaction energy of eq 7 is 0.63 eV. Conversely, it is 0.20 eV in Figure 3b where the energies of CO$_2$ and CO have been corrected. For comparison, such difference is 0.20 eV in experiments (see Table S3. The difference stems from the liquid state of water in eq 7). In terms of the equilibrium potential of the reaction, this all means that PBE predicts it to be at −0.32 V versus RHE, whereas both the correction method and experiments set it at −0.10 V versus RHE. The difference is substantial and amounts to ~220 mV. Note in passing that there are no changes in the energy differences between *COOH and *CO as the corrections are only applied to the gas phase. Although corrections for adsorbates have been proposed before, they escape the subject and scope of this article.

Within the context of CO$_2$RR modeling with the computational hydrogen electrode, the onset potential is given by the largest positive consecutive difference in Figure 3 (U$_\text{onset} = -\max(\Delta G_f, \Delta G_2)/\text{e}^{-}$); see the Computational Methods section). In Figure 3a, such a difference is 0.90 eV, whereas in Figure 3b, it is 0.71 eV so that the predicted onset potentials are −0.90 and −0.71 V versus RHE, respectively. As the experimental value of the onset potential is −0.66 V versus RHE, the deviations from experiments are ~0.24 (as is) and 0.05 V (corrected).

We note that the sizable lowering of the error from 0.24 to 0.05 V is a direct result of correcting gas-phase energetics. To assess whether this is a particularity of Au(111) electrodes or part of a more general trend, we also compared the calculated and experimental onset potentials for Au(100), Au(110), Au$_\text{poly}$, Ag(111), Ag$_\text{poly}$, and Cu$_\text{poly}$. The results in Figure 4a show that DFT data are systematically deviated from the parity line, which results in a MAE of 0.20 V and a MAX of 0.27 V. Conversely, the CO$_2$ and CO corrected data in Figure 4b are located around the parity line with MAE = 0.06 V and MAX = 0.09 V. Substantial improvements are also observed for Au(111) and Au(100) using gas-phase corrections with RPBE (see Figures S7 and S8). Thus, we conclude that models for CO$_2$RR to CO may in general benefit from the gas-phase corrections found in this work.

**CONCLUSIONS**

When interfaces between metals and fluids are simulated at the GGA level, sizable errors may appear in the description of the gas-phase molecules. Here, we proposed a two-step semi-
empirical method to determine gas-phase errors based on the formation energies of 27 different molecules. Furthermore, implementing the corresponding corrections allow for predictions in the analyzed data set of CO2RR and CORR reaction energies that lower by 1 order of magnitude the average and maximum errors with respect to experiments.

The method also shows that the errors for CO2 and CO differ by ~0.4 eV for all the examined exchange–correlation functionals. Thus, an intrinsic limitation of DFT exists for the description of reaction energies containing these two functionals. Therefore, in addition to pinpointing various errors for PBE, PW91, RPBE, and BEEF-vdW exchange–correlation functionals, additional electrocatalysis-related figures and tabulated data, and alternative error analysis using training and extrapolation sets (PDF).

**Notes**

The authors declare no competing financial interest.

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The method also shows that the errors for CO2 and CO differ by ~0.4 eV for all the examined exchange–correlation functionals. Thus, an intrinsic limitation of DFT exists for the description of reaction energies containing these two functionals. Therefore, in addition to pinpointing various errors for PBE, PW91, RPBE, and BEEF-vdW exchange–correlation functionals, additional electrocatalysis-related figures and tabulated data, and alternative error analysis using training and extrapolation sets (PDF).

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